

FBQ1: The two hydrogen nuclei would possibly spin in opposite directions to give

Answer: *Para Hydrogen*

FBQ2: Electronic configuration of Zn is _____

Answer: * $1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10} 4s^2$ *

FBQ3: Period 2 contains elements from lithium to _____

Answer: *Neon*

FBQ4: How many elements are contained in period 4?

Answer: *18*

FBQ5: Period 6 contains 32 elements from _____ to radon

Answer: *Caesium*

FBQ6: The d-block elements are also known as _____

Answer: *Transition*

FBQ7: Both the lanthanides and Actinides are collectively called _____

Answer: *Inner-transition*

FBQ8: The measure of the size of atom is known as _____

Answer: *Atomic radii*

FBQ9: The energy required to remove the least strongly bonded electron from its atom in ground state is known as _____

Answer: *Ionization energy*

FBQ10: The measure of the ability of an atom to accept an electron to form an anion is known as _____

Answer: *Electron affinity*

FBQ11: Atoms of an element with the same atomic number but different mass number is called _____

Answer: *Isotopes*

FBQ12: The mixture of CO and H₂ is known as _____

Answer: *Water gas*

FBQ13: The breaking up of hydrocarbons at higher temperature in the presence of catalyst is called _____

Answer: *Cracking*

FBQ14: Reduction of carbon monoxide with hydrogen gives _____

Answer: *Formaldehyde*

FBQ15: Liquid hydrogen is used as _____

Answer: *Rocket fuel*

FBQ16: The largest use of hydrogen is in the manufacture of _____
Answer: *Ammonia*

FBQ17: Hydrogen combines with a number of elements to form _____
Answer: *Hydrides*

FBQ18: Alkali metals do not occur in free state in nature due to their _____
Answer: *High reactivity*

FBQ19: Potassium is obtained by the reduction of its chloride with _____
Answer: *Sodium vapour*

FBQ20: Lithium and potassium compounds are used in picture tubes of _____
Answer: *Colour television*

FBQ21: The factors which determine the density are atomic weight and _____
Answer: *Volume*

FBQ22: The least electronegative group in the periodic table is _____
Answer: *Alkali metals*

FBQ23: _____ is the most abundant alkali metal in the earth
Answer: *Sodium*

FBQ24: Normal oxides and peroxides of alkali metals are both coloured and _____
Answer: *Diamagnetic*

FBQ25: coloured superoxides of alkali metals are generally _____
Answer: *Paramagnetic*

FBQ26: The phenomenon that occurs when a metal is surrounded by solvent molecules is called _____
Answer: *Solvation*

FBQ27: When a metal is surrounded by water molecule, the phenomenon is known as _____
Answer: *Hydration*

FBQ28: The degree of hydration _____ on moving down the group
Answer: *Decreases*

FBQ29: Which block of the periodic table does group 1 and 2 belong to?
Answer: *S*

FBQ30: The radioactive element in group two is _____

Answer: *Radium*

FBQ31: _____ is used for making atomic fuel containers

Answer: *Beryllium*

FBQ32: The ionization energy of alkaline earth metals _____ on moving down the group

Answer: *Decreases*

FBQ33: All the alkaline earth metals burn in oxygen to form _____

Answer: *Oxides*

FBQ34: All the Group 2 elements except _____ form hydrides by direct combination with hydrogen

Answer: *Beryllium*

FBQ35: Magnesium in chlorophyll is coordinated by how many nitrogen atoms in the heterocyclic porphyrin ring system?

Answer: *4*

Multiple Choice Question (MCQs):

MCQ1: Metallic elements mostly form ____ oxides

Answer: Neutral

MCQ2: Non metallic elements form ____ oxides

Answer: Neutral

MCQ3: The main credit in developing the Periodic Laws goes to ____

Answer: Newlands

MCQ4: According to the _____, no two electrons in the same atom can have the same value of n

Answer: Exclusion Principle

MCQ5: According to Pauli Exclusion Principle, an orbital can have at the most, two electrons of ____ spin

Answer: Equal

MCQ6: The seven f orbital can have ____ number of electrons

Answer: Ten

MCQ7: What is the ground state electronic configuration of silicon? (atomic number 14)

Answer: $1s^2 2s^2 2p^6 3s^2 3p^2$

MCQ8: What is the electronic configuration of Fe^{2+} ?

Answer: $[Ar] 3d^6$

MCQ9: The members of the 4f series are called _____

Answer: Actinides

MCQ10: The bonding within a non metal molecule is largely ____

Answer: Electrovalent

MCQ11: The metallic radius depends to some extent on the crystal structure of the

Answer: Bond

MCQ12: The ionization energies of the noble gases are the highest in the respective periods because

Answer: Stability of electronic configuration

MCQ13: Electron affinity of an atom is a measure of its ability to accept electron to form

Answer: Cation

MCQ14: One of these factors does not affect electron affinity

Answer: Atomic radius

MCQ15: The ability of an atom to attract electrons depends upon the charge on the atom and the ____ of the atom

Answer: Polarity

MCQ16: Across a series of transition elements, the increase in electronegativity is much ____

Answer: Smaller

MCQ17: Hydrogen is the principal element in the solar ____

Answer: System

MCQ18: Hydrogen has three different isotopes having mass numbers

Answer: 1,2,3

MCQ19: How many neutrons does ordinary hydrogen has

Answer: Two

MCQ20: Tritium differs from the other isotopes of hydrogen in being ____

Answer: Reactive

MCQ21: Tritium decays by emission of ____ particle

Answer: Alpha

MCQ22: Two nuclei of hydrogen spin in the same direction to give the form called ____ -

Answer: Para

MCQ23: Hydrogen is manufactured by allowing steam to react with ____ at about 250K

Answer: Coal

MCQ24: Alkaline earth metals are less reactive than alkali metals as they are less ____
Answer: Electronegative

MCQ25: Hydrogen reacts with oxygen at room temperature to form ____
Answer: Water

MCQ26: The process where electrical energy is generated by the reaction of hydrogen and oxygen in the fuel cell is ____
Answer: Cold combustion

MCQ27: Hydrogen bond formed between two atoms of the same molecule is called ____
Answer: Intermolecular

MCQ28: The length of hydrogen bond is much more than the ____
Answer: Electrovalent bond

MCQ29: Group 1 elements are called alkali metals because they form ____ which are strong alkali
Answer: Oxides

MCQ30: Lithium and sodium are extracted by ____ of their fused chlorides
Answer: Oxidation

MCQ31: Lithium in the form of lithium stearate is used for the production of ____
Answer: Lithium Lamp

MCQ32: Potassium chlorate is used in making explosives and ____
Answer: Lubricating oil

MCQ33: Alkali metals form -----
Answer: Divalent cations

MCQ34: Lithium reacts with carbon to form ____ carbides
Answer: Covalent

MCQ35: All the alkali metal nitrates decompose on strong heating to oxides liberating ____
Answer: Carbon IV oxide