

FBQ1: The nature of the charge on Gamma rays is -----charge

Answer: Neutral

FBQ2: The type of charge on Beta ray is -----

Answer: negative

FBQ3: Low ionising power on gases is found on.....ray

Answer: Beta

FBQ4: The nature of the charge on an electron is -----

Answer: Negative

FBQ5: The magnitude of the charge is the same on a proton and an

Answer: Electron

FBQ6: The constituents of the atom are....., electrons and neutrons

Answer: Protons

FBQ7: The relative atomic mass of an element is the mass of one atom of the element compared to (1/12) of the mass of one atom of carbon - 12

Answer: Isotopes

FBQ8: Mass spectrometer is the instrument used to determine fairly accurately the relative atomic masses of

Answer: Elements

FBQ9: When two or more elements combine chemically in fixed proportion by mass, it results to

Answer: compound

FBQ10: Theof compounds and the elements from which they are formed differ from each other

Answer: Properties

FBQ11: Non metals areand gases

Answer: Solids

FBQ12: The process in which two or more light nuclei combine to form a heavier nucleus with a release of energy is known as..... fusion

Answer: Nuclear $\hat{A}$

FBQ13: Electrovalent bonding involves electron transfer from the valence shell of one ...
...to the valence shell of the other

Answer: Atom

FBQ14: The process in which the nucleus of a heavy element is split into two nuclei of nearly equal mass with a release of energy and radiation is known as nuclear

Answer: Fission

FBQ15: The path of deflection of Alpha rays is towards a negative

Answer: Pole

FBQ16: The path on which Gamma rays deflected have no effect onfield

Answer: Electric

FBQ17: The statement that atoms tend to gain or lose electrons until there are eight electrons in their valence shell refers to..... rule_____

Answer: Octet

FBQ18: Michael Faraday's experiment is in.....

Answer: Electricity

FBQ19: Covalent bonding is a form of bonding arises from theof electrons among atoms

Answer: Sharing

FBQ20: Van der Waal's forces are form of bonding exist even between uncombined atoms and nonmolecules?

Answer: Polar

FBQ21: Radioactivity can be associated withnuclei

Answer: Unstable

FBQ22:life is a measure of the time taken for half of the radioactive substance to decay

Answer: Half

FBQ23: A lot of energy is often required to split compounds into theelements

Answer: Constituent

FBQ24: Physical methods can be used to separate awith varying components?

Answer: Mixture

FBQ25: Chemical reactions are explain by combinations and..... of atoms

Answer: Rearrangement

FBQ26: Theory is aproposal to explain an observed statement of facts

Answer: Tested

FBQ27:postulate was discarded as a result of The knowledge of radioactivity.

Answer: Daltons

FBQ28: Faraday,and Millikan conducted experiments that accounted for the evidences about the nature of matter

Answer: Thompson

FBQ29: Chemical symbols and are used to represent elements and compounds

Answer: Formulae

FBQ30: Lustre is a general characteristics of

Answer: Metals

FBQ31: The third particle in the nucleus is.....

Answer: Neutron

FBQ32: Element cannot be

Answer: Split

FBQ33: Substance that can be broken down into elements is called.....

Answer: Compound

FBQ34: Non metalscannot heat

Answer: 8Conduct

FBQ35: Metal is a solid attemperatureÂ

Answer: Room

MCQ1: What is the nature of the charges on an electron?

Answer: negatively

MCQ2: What is the magnitude of the charge on a proton and an electron?

Answer: the same

MCQ3: What is the constituents of the atom?

Answer: protons, electrons and neutrons

MCQ4: For the atom to be electrically neutral, what should be the magnitude of the charges of the protons and electrons?

Answer: number of protons and electrons must be equalÂ

MCQ5: What is the relative atomic mass of an element?Â

Answer: the mass of one atom of the element compared to (1/12) of the mass of one atom of carbon - 12 isotopes

MCQ6: What instrument would you use to determine fairly accurately the relative atomic masses of elements

Answer: mass spectrometer

MCQ7: What do Chemical symbols and formulae used to represent ?

Answer: elements and compounds

MCQ8: Which of these laws was discarded as a result of The knowledge of radioactivity?

Answer: Daltons postulate

MCQ9: Which of these can be associated with unstable nuclei?

Answer: Radioactivity

MCQ10: Which of these statements best define the phenomenon of half-life?

Answer: a measure of the time taken for half of the radioactive substance to decay

MCQ11: What radiations does a radioactive element in its decay emit?

Answer: Alpha, beta and gamma rays

MCQ12: What type of charges do Alpha rays bear?

Answer: positively charged

MCQ13: Identify the correct statement about Alpha rays

Answer: high ionising power on gases

MCQ14: What is the nature of the charge on Gamma rays?

Answer: neutrally charged

MCQ15: What type of charges do Beta rays carry?

Answer: negatively charged

MCQ16: Which of these statements hold true about Beta rays?

Answer: low ionising power on gases

MCQ17: Which of these statements hold true about Gamma rays?

Answer: very low ionising power on gases

MCQ18: Which of these statements hold true for an element ?

Answer: cannot be split

MCQ19: What is a pure substance that can be broken down into elements called

Answer: Compound

MCQ20: How many groups can you subdivide an element into?

Answer: 2

MCQ21: What results when two or more elements combine chemically in fixed proportion by mass.

Answer: compound

MCQ22: Which of these can be associated with unstable nuclei?

Answer: Radioactivity

MCQ23: How much energy is often required to split compounds into the constituent elements

Answer: a lot of

MCQ24: What is the numbers of compounds available.

Answer: Limitless

MCQ25: How would you separate a mixture with varying components?

Answer: physical methods

MCQ26: Which of these methods can be used to separate a mixture

Answer: all of the options

MCQ27: How would you explain chemical reactions of atoms

Answer: combinations and rearrangement

MCQ28: What is a tested proposal to explain an observed statement of fact

Answer: Theory

MCQ29: Which law states that all elements are made up of small, indivisible particles called atoms

Answer: Daltons Atomic theory

MCQ30: which of these scientists conducted experiments that accounted for the evidences about the nature of matter?

Answer: Faraday, Thompson and Millikan

MCQ31: What form of bonding exist even between uncombined atoms and non polar molecules?

Answer: Vaander Waal's forces

MCQ32: Which of these is the nature of Matter?

Answer: Electrical

MCQ33: The process in which two or more light nuclei combine to form a heavier nucleus with a release of energy is known as

Answer: nuclear fussion

MCQ34: Which of these is used for radioactivity?

Answer: all of the options

MCQ35: What does the statement that atoms tend to gain or lose electrons until there are eight electrons in their valence shell refers to?

Answer: Covalent bonding